

Nomenclature Worksheet 5: Ionic Compounds Summary

Name the following compounds:

1. CaF_2 _____
2. Na_2O _____
3. BaS _____
4. CuSO_4 _____
5. Fe_2O_3 _____
6. HgCl_2 _____
7. AgNO_3 _____
8. MgCO_3 _____
9. $\text{KC}_2\text{H}_3\text{O}_2$ _____
10. $\text{K}_2\text{Cr}_2\text{O}_7$ _____
11. $\text{Al}(\text{OH})_3$ _____
12. PbBr_2 _____
13. ZnSO_3 _____
14. NaHCO_3 _____
15. NH_4Cl _____
16. Li_3PO_4 _____
17. SnCl_2 _____
18. $\text{Al}(\text{NO}_2)_3$ _____
19. Rb_2CrO_4 _____
20. KMnO_4 _____
21. CuCl _____
22. FeSO_4 _____

Give the formula for each compound:

23. sodium fluoride _____
24. potassium sulfide _____
25. calcium carbonate _____
26. magnesium hydroxide _____
27. zinc nitrate _____
28. silver acetate _____
29. copper (II) oxide _____
30. iron (III) chloride _____
31. barium chromate _____
32. aluminum oxide _____
33. lead (II) sulfate _____
34. tin (IV) oxalate _____
35. calcium phosphate _____
36. lithium permanganate _____
37. mercury (I) nitrate _____
38. radium sulfite _____
39. chromium (III) chloride _____
40. ammonium sulfide _____
41. copper (II) acetate _____
42. calcium bicarbonate _____
43. tin (II) oxide _____
44. silver sulfite _____

Nomenclature Worksheet 6: Binary Covalent Compounds

Please complete the following table:

Name of <i>Covalent</i> Compound	Formula of <i>Covalent</i> Compound
1. carbon dioxide	
2. phosphorus triiodide	
3. sulfur dichloride	
4. nitrogen trifluoride	
5. dioxygen difluoride	
	6. N ₂ F ₄
	7. SCl ₄
	8. ClF ₃
	9. SiO ₂
	10. P ₄ O ₁₀

Determine whether the following compounds are **covalent** or **ionic** and give them their proper names.

1. Ba(NO₃)₂
2. CO
3. PCl₃
4. KI
5. CF₄
6. MgO
7. Cu₂S
8. SO₂
9. NCl₃
10. XeF₆