Set A: Metal - acid reaction

Objective: To test your ability to determine which metal - acid reaction will occur, and be able to write a complete balanced equations

Below, equation showing reaction of a metal with an acid is given.

If the reaction will occur, write the products (salt and hydrogen gas) formulas to the right of the arrow.

Be sure to balance these equations.

If the reaction will not occur, write "no reaction" to the right of the arrow.

USF TABLE I

3 HZ

"	if the reaction will not occur, write			Te "no reaction" to the right of the arrow.
1.	<i>Metal</i> Ca	<i>†</i> +	Acid 2HCI	Salt + hydrogen gas> Cacl2 H 2
2.	Ag	+	H N O₃	> NA
3.	(∙ Li	+	2H₃PO₄	> 2 Li3Poy + 3Hz
4.	Mg	+	HC₂H₃O₂	> Mg ((2H302)2 + HZ
5.	Zn	+	2.HCl	> ZnCl2 + Hz.
6.	Au	+	H ₂ SO ₄	> No reaction
_	^ .			•

Set B: Neutralization reaction

Objective: To test your ability to complete and balance neutralization reactions.

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Below, incomplete neutralization equation are given.

Write the products of each reaction to the right of the arrow. Be sure to balance the equations .

1					•	
8.	Acid HCl	+ Base + KOH	e 1	salt > KC		Water
9.	HNO₃	+ NaO		> NaNO2		H20 H20
10.	H ₂ SO ₄	+ 2 кон		> K2soy		2 H20
11.	H ₂ CO ₃	+ Ca(O)H) ₂	> Ca (03	+	2420
12.	H ₃ PO ₄	+ 3LiOH	>	Liz Poy	+ 3	3H20
13.	ZHC2H3O2	+ Ba(OI	H) ₂ >	Ba((2H302)	2 + 6	2 HZO