

Titration Practice Worksheet

Find the requested quantities in the following problems:

- 1) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?

0.043M HCl

- 2) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

0.0036 M NaOH

3)

- If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulfuric acid solution (H_2SO_4), what is the concentration of the H_2SO_4 solution?

1 M H_2SO_4

4)

- Can I titrate a solution of unknown concentration with another solution of unknown concentration and still get a meaningful answer? Explain your answer in a few sentences.

NO

~~5)~~

- Explain the difference between an endpoint and equivalence point in a titration.

↓
Stop
titration

↓
when
neutralized