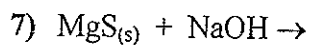
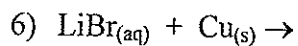
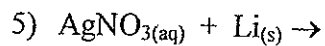
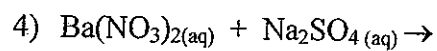
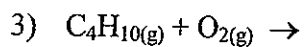
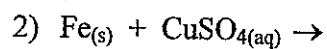
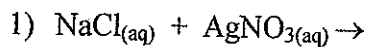
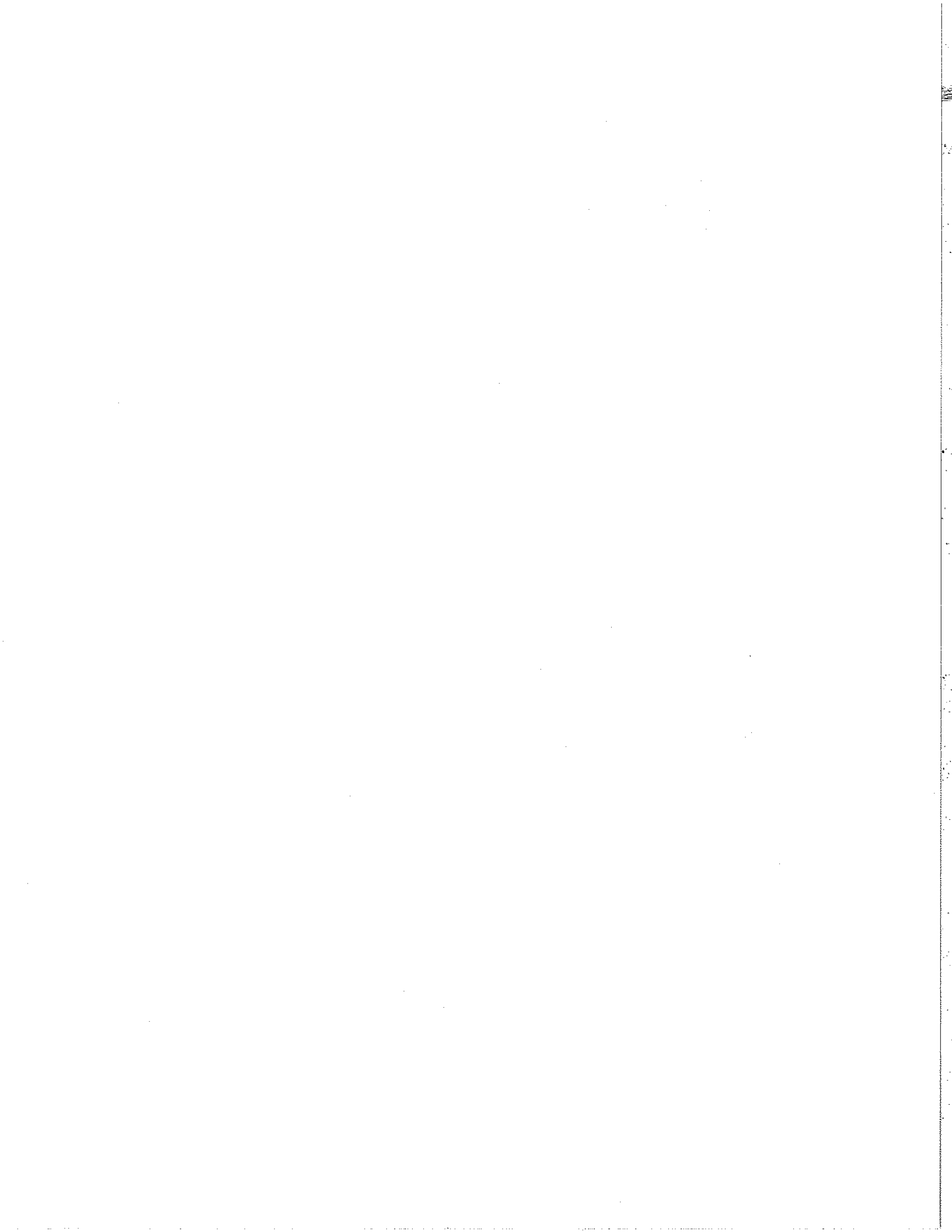


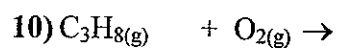
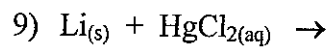
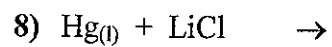
### : Mixed Types of Reactions

Directions: For each of the following reactions:

- 1) Identify the type of reaction (Single Replacement, combination, double replacement, decomposition, or combustion)
- 2) Complete the reaction (Put NR if no reaction takes place)
- 3) Balance it
- 4) Indicate states of matter for each substance (solid, liquid, gas, aqueous)







11) Ammonium Nitrate is added to sodium chloride

12) Lithium Chloride is added to Zinc Phosphate

13) Zinc is added to lithium chloride

14) Iron is added to a solution of silver nitrate

15) A solution of copper II sulfate is added to an iron nail

16) Octane ( $C_8H_{18}$ ) is burned in air

17) A solution of Tin IV sulfate is added to a solution of ammonium hydroxide

18) Sodium hydroxide is added to hydrochloric acid

19) Calcium hydroxide is added to sulfuric acid

20) Nickel is added to hydrochloric acid

21) Strontium is added to water

22) Hydrochloric acid is added to copper metal

23) Solid bismuth is added to a solution of barium hydroxide.

24) Methanol ( $\text{CH}_3\text{OH}$ ) is burned in air

25) Solid gold is added to hydrochloric acid

