Balance the following equations + Label type of reaction

1) 
$$_{\text{HgO}}$$
  $\rightarrow$   $_{\text{Hg}}$  +  $_{\text{O}}$ 

2) 
$$\_HCl + \_Mg \rightarrow \_H_2 + MgCl_2$$

3) 
$$\_CH_4 + \_O_2 \rightarrow \_CO_2 + \_H_2O$$

4) 
$$C_6H_{12}O_6 + O_2 \rightarrow CO_2 + H_2O$$

5) 
$$\_H_2 + \_O_2 \rightarrow \_H_2O$$

6) 
$$\_H_2 + \_N_2 \rightarrow \_NH_3$$

7) 
$$\_NO + \_O_2 \rightarrow \_NO_2$$

8) 
$$\_Al_2O_3$$
  $\rightarrow$   $\_Al +  $\_O_2$$ 

9) 
$$\_CaO+$$
  $\_H_2O \rightarrow$   $\_Ca(OH)_2$ 

10) Hydrogen gas reacts with iodine to produce hydroiodic acid.

11) Sulfur reacts with oxygen to produce sulfur dioxide.

12) Calcium acetate reacts with sodium carbonate to produce calcium carbonate and sodium acetate.
13) Iron combines with oxygen and water to form iron (III) hydroxide
14) Sulfur trioxide is bubbled through water to produce sulfuric acid
15) Copper-bottomed cooking pans turn black because copper combines with oxyger to form copper (II) oxide.
16) Magnesium hydroxide neutralizes stomach acid, HCl, to produce magnesium chloride and water.