

C

Reference Tables for Physical Setting/CHEMISTRY

Table A
Standard Temperature and Pressure

Name	Value	Unit
Standard Pressure	101.3 kPa 1 atm	kilopascal atmosphere
Standard Temperature	273 K 0°C	kelvin degree Celsius

Table B
Physical Constants for Water

Heat of Fusion	334 J/g
Heat of Vaporization	2260 J/g
Specific Heat Capacity of H ₂ O(ℓ)	4.18 J/g•K

Table C
Selected Prefixes

Factor	Prefix	Symbol
10 ³	kilo-	k
10 ⁻¹	deci-	d
10 ⁻²	centi-	c
10 ⁻³	milli-	m
10 ⁻⁶	micro-	μ
10 ⁻⁹	nano-	n
10 ⁻¹²	pico-	p

Table D
Selected Units

Symbol	Name	Quantity
m	meter	length
g	gram	mass
Pa	pascal	pressure
K	kelvin	temperature
mol	mole	amount of substance
J	joule	energy, work, quantity of heat
s	second	time
min	minute	time
h	hour	time
d	day	time
y	year	time
L	liter	volume
ppm	parts per million	concentration
M	molarity	solution concentration
u	atomic mass unit	atomic mass

Table E
Selected Polyatomic Ions

Formula	Name	Formula	Name
H_3O^+	hydronium	CrO_4^{2-}	chromate
Hg_2^{2+}	mercury(I)	$\text{Cr}_2\text{O}_7^{2-}$	dichromate
NH_4^+	ammonium	MnO_4^-	permanganate
$\left. \begin{array}{l} \text{C}_2\text{H}_3\text{O}_2^- \\ \text{CH}_3\text{COO}^- \end{array} \right\}$	acetate	NO_2^-	nitrite
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	hydrogen carbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_4^{3-}	phosphate
ClO^-	hypochlorite	SCN^-	thiocyanate
ClO_2^-	chlorite	SO_3^{2-}	sulfite
ClO_3^-	chlorate	SO_4^{2-}	sulfate
ClO_4^-	perchlorate	HSO_4^-	hydrogen sulfate
		$\text{S}_2\text{O}_3^{2-}$	thiosulfate

Table F
Solubility Guidelines for Aqueous Solutions

Ions That Form Soluble Compounds	Exceptions	Ions That Form Insoluble Compounds*	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		carbonate (CO_3^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Mg^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
acetate ($\text{C}_2\text{H}_3\text{O}_2^-$ or CH_3COO^-)		sulfide (S^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogen carbonate (HCO_3^-)		hydroxide (OH^-)	when combined with Group 1 ions, Ca^{2+} , Ba^{2+} , Sr^{2+} , or ammonium (NH_4^+)
chlorate (ClO_3^-)			
halides (Cl^- , Br^- , I^-)	when combined with Ag^+ , Pb^{2+} , or Hg_2^{2+}		
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*compounds having very low solubility in H_2O

Table G
Solubility Curves at Standard Pressure

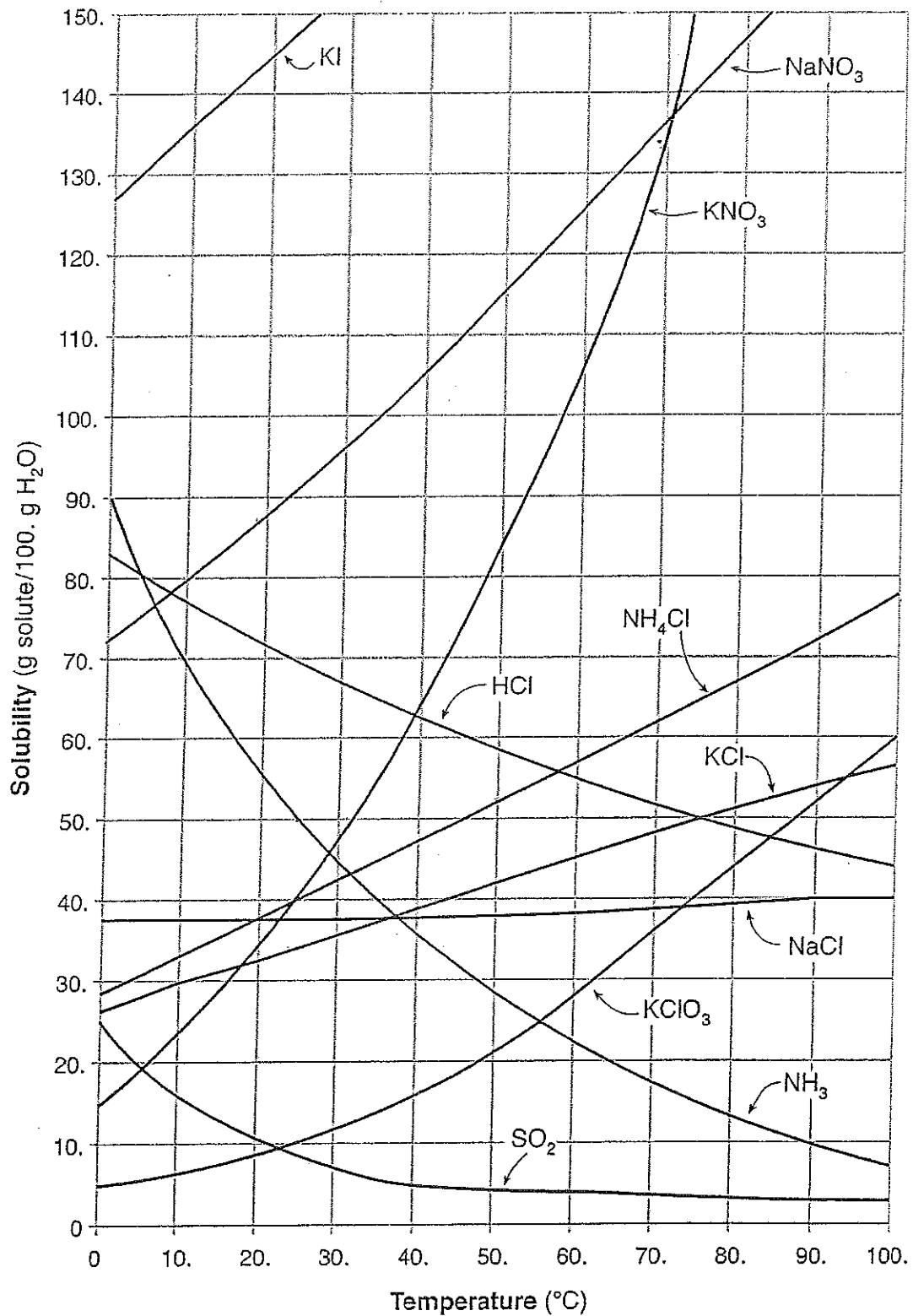


Table H
Vapor Pressure of Four Liquids

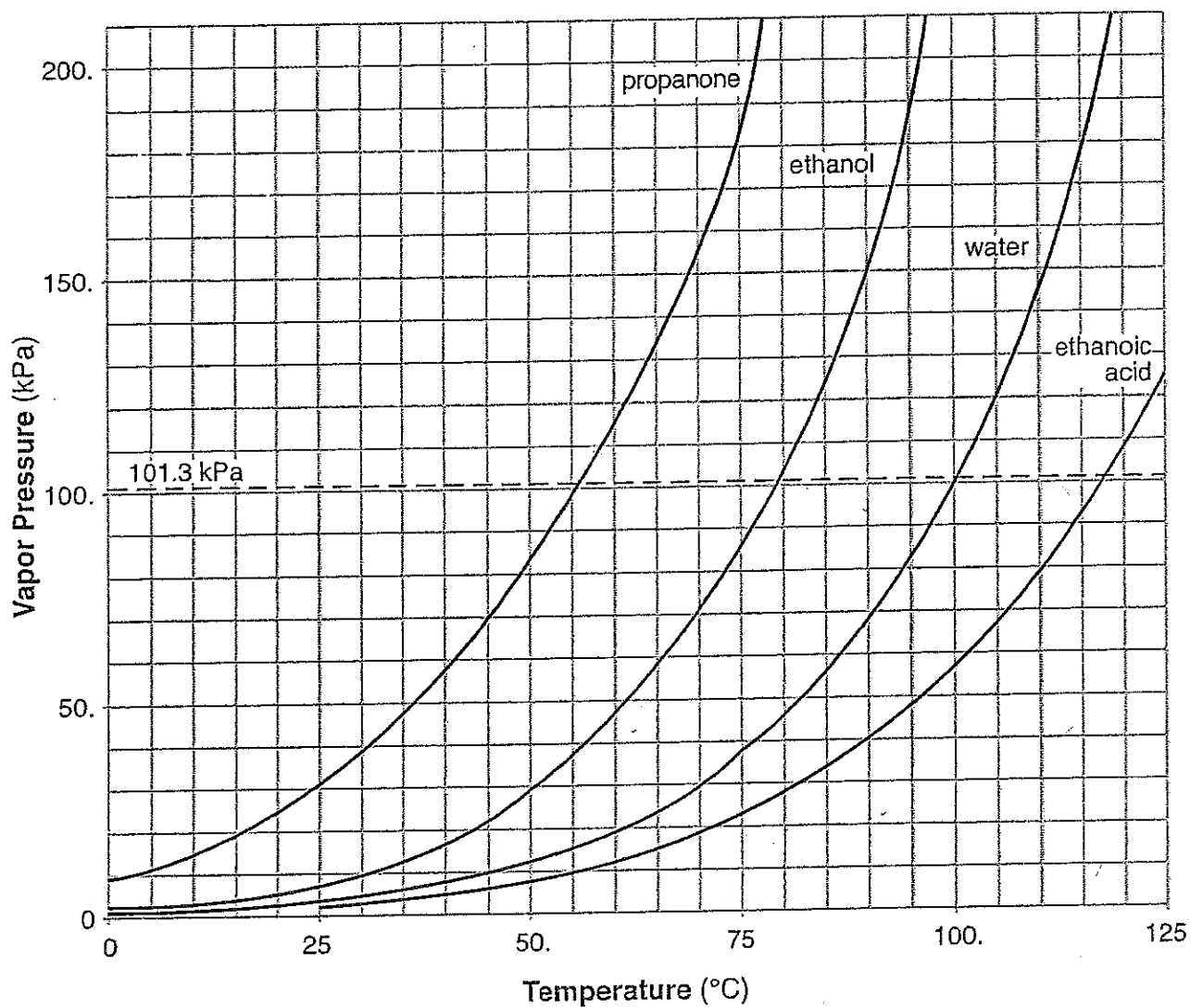


Table I
Heats of Reaction at 101.3 kPa and 298 K

Reaction	ΔH (kJ)*
$\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \longrightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\ell)$	-890.4
$\text{C}_3\text{H}_8(\text{g}) + 5\text{O}_2(\text{g}) \longrightarrow 3\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\ell)$	-2219.2
$2\text{C}_8\text{H}_{18}(\ell) + 25\text{O}_2(\text{g}) \longrightarrow 16\text{CO}_2(\text{g}) + 18\text{H}_2\text{O}(\ell)$	-10943
$2\text{CH}_3\text{OH}(\ell) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\ell)$	-1452
$\text{C}_2\text{H}_5\text{OH}(\ell) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\ell)$	-1367
$\text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + 6\text{O}_2(\text{g}) \longrightarrow 6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\ell)$	-2804
$2\text{CO}(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{CO}_2(\text{g})$	-566.0
$\text{C}(\text{s}) + \text{O}_2(\text{g}) \longrightarrow \text{CO}_2(\text{g})$	-393.5
$4\text{Al}(\text{s}) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{Al}_2\text{O}_3(\text{s})$	-3351
$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{NO}(\text{g})$	+182.6
$\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \longrightarrow 2\text{NO}_2(\text{g})$	+66.4
$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\text{g})$	-483.6
$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\ell)$	-571.6
$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$	-91.8
$2\text{C}(\text{s}) + 3\text{H}_2(\text{g}) \longrightarrow \text{C}_2\text{H}_6(\text{g})$	-84.0
$2\text{C}(\text{s}) + 2\text{H}_2(\text{g}) \longrightarrow \text{C}_2\text{H}_4(\text{g})$	+52.4
$2\text{C}(\text{s}) + \text{H}_2(\text{g}) \longrightarrow \text{C}_2\text{H}_2(\text{g})$	+227.4
$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \longrightarrow 2\text{HI}(\text{g})$	+53.0
$\text{KNO}_3(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$	+34.89
$\text{NaOH}(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$	-44.51
$\text{NH}_4\text{Cl}(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$	+14.78
$\text{NH}_4\text{NO}_3(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{NH}_4^+(\text{aq}) + \text{NO}_3^-(\text{aq})$	+25.69
$\text{NaCl}(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$	+3.88
$\text{LiBr}(\text{s}) \xrightarrow{\text{H}_2\text{O}} \text{Li}^+(\text{aq}) + \text{Br}^-(\text{aq})$	-48.83
$\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \longrightarrow \text{H}_2\text{O}(\ell)$	-55.8

*The ΔH values are based on molar quantities represented in the equations. A minus sign indicates an exothermic reaction.

Table J
Activity Series**

Most Active	Metals	Nonmetals	Most Active
↓	Li	F_2	↓
	Rb	Cl_2	
	K	Br_2	
	Cs	I_2	
	Ba -		
	Sr		
	Ca		
	Na		
	Mg		
	Al		
	Ti		
	Mn		
	Zn		
	Cr		
	Fe -		
	Co		
	Ni -		
	Sn		
	Pb		
H_2			
Cu -			
Ag			
Au			
Least Active			Least Active

**Activity Series is based on the hydrogen standard. H_2 is not a metal.

Table K
Common Acids

Formula	Name
HCl(aq)	hydrochloric acid
HNO ₂ (aq)	nitrous acid
HNO ₃ (aq)	nitric acid
H ₂ SO ₃ (aq)	sulfurous acid
H ₂ SO ₄ (aq)	sulfuric acid
H ₃ PO ₄ (aq)	phosphoric acid
H ₂ CO ₃ (aq) or CO ₂ (aq)	carbonic acid
CH ₃ COOH(aq) or HC ₂ H ₃ O ₂ (aq)	ethanoic acid (acetic acid)

Table L
Common Bases

Formula	Name
NaOH(aq)	sodium hydroxide
KOH(aq)	potassium hydroxide
Ca(OH) ₂ (aq)	calcium hydroxide
NH ₃ (aq)	aqueous ammonia

Table M
Common Acid–Base Indicators

Indicator	Approximate pH Range for Color Change	Color Change
methyl orange	3.1–4.4	red to yellow
bromthymol blue	6.0–7.6	yellow to blue
phenolphthalein	8–9	colorless to pink
litmus	4.5–8.3	red to blue
bromocresol green	3.8–5.4	yellow to blue
thymol blue	8.0–9.6	yellow to blue

Source: *The Merck Index*, 14th ed., 2006, Merck Publishing Group

Table N
Selected Radioisotopes

Nuclide	Half-Life	Decay Mode	Nuclide Name
¹⁹⁸ Au	2.695 d	β ⁻	gold-198
¹⁴ C	5715 y	β ⁻	carbon-14
³⁷ Ca	182 ms	β ⁺	calcium-37
⁶⁰ Co	5.271 y	β ⁻	cobalt-60
¹³⁷ Cs	30.2 y	β ⁻	cesium-137
⁵³ Fe	8.51 min	β ⁺	iron-53
²²⁰ Fr	27.4 s	α	francium-220
³ H	12.31 y	β ⁻	hydrogen-3
¹³¹ I	8.021 d	β ⁻	iodine-131
³⁷ K	1.23 s	β ⁺	potassium-37
⁴² K	12.36 h	β ⁻	potassium-42
⁸⁵ Kr	10.73 y	β ⁻	krypton-85
¹⁶ N	7.13 s	β ⁻	nitrogen-16
¹⁹ Ne	17.22 s	β ⁺	neon
³² P	14.28 d	β ⁻	phosphorus-32
²³⁹ Pu	2.410 × 10 ⁴ y	α	plutonium-239
²²⁶ Ra	1599 y	α	radium-226
²²² Rn	3.823 d	α	radon-222
⁹⁰ Sr	29.1 y	β ⁻	strontium-90
⁹⁹ Tc	2.13 × 10 ⁵ y	β ⁻	technetium-99
²³² Th	1.40 × 10 ¹⁰ y	α	thorium-232
²³³ U	1.592 × 10 ⁵ y	α	uranium-233
²³⁵ U	7.04 × 10 ⁸ y	α	uranium-235
²³⁸ U	4.47 × 10 ⁹ y	α	uranium-238

Source: *CRC Handbook of Chemistry and Physics*, 91st ed., 2010–2011, CRC Press

Table O
Symbols Used in Nuclear Chemistry

Name	Notation	Symbol
alpha particle	${}^4_2\text{He}$ or ${}^4_2\alpha$	α
beta particle	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	β^-
gamma radiation	${}^0_0\gamma$	γ
neutron	${}^1_0\text{n}$	n
proton	${}^1_1\text{H}$ or ${}^1_1\text{p}$	p
positron	${}^0_{+1}\text{e}$ or ${}^0_{+1}\beta$	β^+

Table P
Organic Prefixes

Prefix	Number of Carbon Atoms
meth-	1
eth-	2
prop-	3
but-	4
pent-	5
hex-	6
hept-	7
oct-	8
non-	9
dec-	10

Table Q
Homologous Series of Hydrocarbons

Name	General Formula	Examples	
		Name	Structural Formula
alkanes	$\text{C}_n\text{H}_{2n+2}$	ethane	$\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
alkenes	C_nH_{2n}	ethene	$\begin{array}{c} \text{H} \quad \quad \text{H} \\ \diagdown \quad \diagup \\ \text{C}=\text{C} \\ \diagup \quad \diagdown \\ \text{H} \quad \quad \text{H} \end{array}$
alkynes	$\text{C}_n\text{H}_{2n-2}$	ethyne	$\text{H}-\text{C}\equiv\text{C}-\text{H}$

Note: n = number of carbon atoms

Table R
Organic Functional Groups

Class of Compound	Functional Group	General Formula	Example
halide (halocarbon)	-F (fluoro-) -Cl (chloro-) -Br (bromo-) -I (iodo-)	$R-X$ (X represents any halogen)	$CH_3CHClCH_3$ 2-chloropropane
alcohol	-OH	$R-OH$	$CH_3CH_2CH_2OH$ 1-propanol
ether	-O-	$R-O-R'$	$CH_3OCH_2CH_3$ methyl ethyl ether
aldehyde	$\begin{array}{c} O \\ \\ -C-H \end{array}$	$\begin{array}{c} O \\ \\ R-C-H \end{array}$	$\begin{array}{c} O \\ \\ CH_3CH_2C-H \end{array}$ propanal
ketone	$\begin{array}{c} O \\ \\ -C- \end{array}$	$\begin{array}{c} O \\ \\ R-C-R' \end{array}$	$\begin{array}{c} O \\ \\ CH_3CCH_2CH_2CH_3 \end{array}$ 2-pentanone
organic acid	$\begin{array}{c} O \\ \\ -C-OH \end{array}$	$\begin{array}{c} O \\ \\ R-C-OH \end{array}$	$\begin{array}{c} O \\ \\ CH_3CH_2C-OH \end{array}$ propanoic acid
ester	$\begin{array}{c} O \\ \\ -C-O- \end{array}$	$\begin{array}{c} O \\ \\ R-C-O-R' \end{array}$	$\begin{array}{c} O \\ \\ CH_3CH_2COCH_3 \end{array}$ methyl propanoate
amine	$\begin{array}{c} \\ -N- \end{array}$	$\begin{array}{c} R' \\ \\ R-N-R'' \end{array}$	$CH_3CH_2CH_2NH_2$ 1-propanamine
amide	$\begin{array}{c} O \\ \\ -C-NH \end{array}$	$\begin{array}{c} O \quad R' \\ \quad \\ R-C-NH \end{array}$	$\begin{array}{c} O \\ \\ CH_3CH_2C-NH_2 \end{array}$ propanamide

Note: R represents a bonded atom or group of atoms.

Periodic Table of the Elements

Period	1	2
1	1.00794 1 H	4.00260 2 He

18	4.00260 2 He
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KEY

Atomic Mass → 12.011

Symbol → **C**

Atomic Number → 6

Electron Configuration → 2-4

Selected Oxidation States → -4, +2, +4

Relative atomic masses are based on $^{12}\text{C} = 12$ (exact)

Note: Numbers in parentheses are mass numbers of the most stable or common isotope.

	Group 1		Group 2										Group 3										Group 4										Group 5										Group 6										Group 7										Group 8										Group 9										Group 10										Group 11										Group 12										Group 13										Group 14										Group 15										Group 16										Group 17										Group 18									
1	6.941 3 2-1 Li	9.01218 4 2-2 Be	22.98977 11 2-8-1 Na	24.305 12 2-8-2 Mg	39.0983 19 2-8-1 K	40.08 20 2-8-2 Ca	44.9559 21 2-8-2 Sc	47.867 22 2-8-10-2 Ti	50.9415 23 2-8-11-2 V	51.966 24 2-8-13-1 Cr	54.938 25 2-8-13-2 Mn	55.845 26 2-8-14-2 Fe	58.932 27 2-8-15-2 Co	58.933 28 2-8-16-2 Ni	63.546 29 2-8-18-1 Cu	65.409 30 2-8-18-2 Zn	69.723 31 2-8-18-3 Ga	72.64 32 2-8-18-4 Ge	74.9216 33 2-8-18-5 As	78.96 34 2-8-18-6 Se	79.904 35 2-8-18-7 Br	83.904 36 2-8-18-8 Kr	85.468 37 2-8-18-9 Rb	87.62 38 2-8-18-10 Sr	91.224 39 2-8-18-9-2 Y	92.9064 40 2-8-18-10-2 Zr	95.94 41 2-8-18-12-1 Nb	97.907 42 2-8-18-13-2 Mo	101.07 43 2-8-18-15-1 Tc	102.906 44 2-8-18-16-1 Ru	105.42 45 2-8-18-16-2 Rh	106.42 46 2-8-18-16-3 Pd	107.868 47 2-8-18-16-4 Ag	112.41 48 2-8-18-18-3 Cd	114.818 49 2-8-18-18-4 In	118.71 50 2-8-18-18-5 Sn	121.760 51 2-8-18-18-6 Sb	127.60 52 2-8-18-18-7 Te	126.905 53 2-8-18-18-8 I	127.60 54 2-8-18-18-9 Xe	132.905 55 2-8-18-18-10 Cs	137.33 56 2-8-18-18-11-2 Ba	138.905 57 2-8-18-18-9-2 La	178.49 58 2-8-18-18-10-2 Ce	178.49 59 2-8-18-18-11-2 Pr	180.948 60 2-8-18-18-12-2 Nd	183.84 61 2-8-18-18-13-2 Pm	186.207 62 2-8-18-18-14-2 Sm	187.037 63 2-8-18-18-14-2 Eu	188.906 64 2-8-18-18-15-2 Gd	192.223 65 2-8-18-18-16-2 Tb	193.224 66 2-8-18-18-17-1 Dy	195.08 67 2-8-18-18-17-1 Ho	196.967 68 2-8-18-18-18-1 Er	200.59 69 2-8-18-18-19-3 Tm	204.383 70 2-8-18-18-19-3 Yb	207.2 71 2-8-18-18-20-3 Lu	208.980 72 2-8-18-18-20-3 Hf	208.980 73 2-8-18-18-20-3 Ta	208.980 74 2-8-18-18-20-3 W	208.980 75 2-8-18-18-20-3 Re	208.980 76 2-8-18-18-20-3 Os	208.980 77 2-8-18-18-20-3 Ir	208.980 78 2-8-18-18-20-3 Pt	208.980 79 2-8-18-18-20-3 Au	208.980 80 2-8-18-18-20-3 Hg	208.980 81 2-8-18-18-20-3 Tl	208.980 82 2-8-18-18-20-3 Pb	208.980 83 2-8-18-18-20-3 Bi	208.980 84 2-8-18-18-20-3 Po	208.980 85 2-8-18-18-20-3 At	208.980 86 2-8-18-18-20-3 Rn	208.980 87 2-8-18-18-20-3 Fr	208.980 88 2-8-18-18-20-3 Ra	208.980 89 2-8-18-18-20-3 Ac	208.980 90 2-8-18-18-20-3 Th	208.980 91 2-8-18-18-20-3 Pa	208.980 92 2-8-18-18-20-3 U	208.980 93 2-8-18-18-20-3 Np	208.980 94 2-8-18-18-20-3 Pu	208.980 95 2-8-18-18-20-3 Am	208.980 96 2-8-18-18-20-3 Cm	208.980 97 2-8-18-18-20-3 Bk	208.980 98 2-8-18-18-20-3 Cf	208.980 99 2-8-18-18-20-3 Es	208.980 100 2-8-18-18-20-3 Fm	208.980 101 2-8-18-18-20-3 Md	208.980 102 2-8-18-18-20-3 No	208.980 103 2-8-18-18-20-3 Lr	208.980 104 2-8-18-18-20-3 Rf	208.980 105 2-8-18-18-20-3 Db	208.980 106 2-8-18-18-20-3 Sg	208.980 107 2-8-18-18-20-3 Bh	208.980 108 2-8-18-18-20-3 Hs	208.980 109 2-8-18-18-20-3 Mt	208.980 110 2-8-18-18-20-3 Ds	208.980 111 2-8-18-18-20-3 Rg	208.980 112 2-8-18-18-20-3 Cn	208.980 113 2-8-18-18-20-3 Nh	208.980 114 2-8-18-18-20-3 Fl	208.980 115 2-8-18-18-20-3 Mc	208.980 116 2-8-18-18-20-3 Lv	208.980 117 2-8-18-18-20-3 Ts	208.980 118 2-8-18-18-20-3 Og																																																																				

140.116 58 Ce	140.908 59 Pr	144.24 60 Nd	145 61 Pm	150.36 62 Sm	151.964 63 Eu	157.25 64 Gd	158.925 65 Tb	162.500 66 Dy	164.930 67 Ho	167.259 68 Er	168.934 69 Tm	173.04 70 Yb	174.967 71 Lu
232.038 90 Th	231.036 91 Pa	238.029 92 U	237 93 Np	244 94 Pu	243 95 Am	247 96 Cm	247 97 Bk	251 98 Cf	252 99 Es	257 100 Fm	258 101 Md	259 102 No	262 103 Lr

*denotes the presence of (2-8-) for elements 72 and above

**The systematic names and symbols for elements of atomic numbers 113 and above will be used until the approval of trivial names by IUPAC.

Table S
Properties of Selected Elements

Atomic Number	Symbol	Name	First Ionization Energy (kJ/mol)	Electro-negativity	Melting Point (K)	Boiling* Point (K)	Density** (g/cm ³)	Atomic Radius (pm)
1	H	hydrogen	1312	2.2	14	20.	0.000082	32
2	He	helium	2372	—	—	4	0.000164	37
3	Li	lithium	520.	1.0	454	1615	0.534	130.
4	Be	beryllium	900.	1.6	1560.	2744	1.85	99
5	B	boron	801	2.0	2348	4273	2.34	84
6	C	carbon	1086	2.6	—	—	—	75
7	N	nitrogen	1402	3.0	63	77	0.001145	71
8	O	oxygen	1314	3.4	54	90.	0.001308	64
9	F	fluorine	1681	4.0	53	85	0.001553	60.
10	Ne	neon	2081	—	24	27	0.000825	62
11	Na	sodium	496	0.9	371	1156	0.97	160.
12	Mg	magnesium	738	1.3	923	1363	1.74	140.
13	Al	aluminum	578	1.6	933	2792	2.70	124
14	Si	silicon	787	1.9	1687	3538	2.3296	114
15	P	phosphorus (white)	1012	2.2	317	554	1.823	109
16	S	sulfur (monoclinic)	1000.	2.6	388	718	2.00	104
17	Cl	chlorine	1251	3.2	172	239	0.002898	100.
18	Ar	argon	1521	—	84	87	0.001633	101
19	K	potassium	419	0.8	337	1032	0.89	200.
20	Ca	calcium	590.	1.0	1115	1757	1.54	174
21	Sc	scandium	633	1.4	1814	3109	2.99	159
22	Ti	titanium	659	1.5	1941	3560.	4.506	148
23	V	vanadium	651	1.6	2183	3680.	6.0	144
24	Cr	chromium	653	1.7	2180.	2944	7.15	130.
25	Mn	manganese	717	1.6	1519	2334	7.3	129
26	Fe	iron	762	1.8	1811	3134	7.87	124
27	Co	cobalt	760.	1.9	1768	3200.	8.86	118
28	Ni	nickel	737	1.9	1728	3186	8.90	117
29	Cu	copper	745	1.9	1358	2835	8.96	122
30	Zn	zinc	906	1.7	693	1180.	7.134	120.
31	Ga	gallium	579	1.8	303	2477	5.91	123
32	Ce	germanium	762	2.0	1211	3106	5.3234	120.
33	As	arsenic (gray)	944	2.2	1090.	—	5.75	120.
34	Se	selenium (gray)	941	2.6	494	958	4.809	118
35	Br	bromine	1140.	3.0	266	332	3.1028	117
36	Kr	krypton	1351	—	116	120.	0.003425	116
37	Rb	rubidium	403	0.8	312	961	1.53	215
38	Sr	strontium	549	1.0	1050.	1655	2.64	190.
39	Y	yttrium	600.	1.2	1795	3618	4.47	176
40	Zr	zirconium	640.	1.3	2128	4682	6.52	164

Atomic Number	Symbol	Name	First Ionization Energy (kJ/mol)	Electro-negativity	Melting Point (K)	Boiling Point (K)	Density** (g/cm ³)	Atomic Radius (pm)
41	Nb	niobium	652	1.6	2750.	5017	8.57	156
42	Mo	molybdenum	684	2.2	2896	4912	10.2	146
43	Tc	technetium	702	2.1	2430.	4538	11	138
44	Ru	ruthenium	710.	2.2	2606	4423	12.1	136
45	Rh	rhodium	720.	2.3	2237	3968	12.4	134
46	Pd	palladium	804	2.2	1828	3236	12.0	130.
47	Ag	silver	731	1.9	1235	2435	10.5	136
48	Cd	cadmium	868	1.7	594	1040.	8.69	140.
49	In	indium	558	1.8	430.	2345	7.31	142
50	Sn	tin (white)	709	2.0	505	2875	7.287	140.
51	Sb	antimony (gray)	831	2.1	904	1860.	6.68	140.
52	Te	tellurium	869	2.1	723	1261	6.232	137
53	I	iodine	1008	2.7	387	457	4.933	136
54	Xe	xenon	1170.	2.6	161	165	0.005366	136
55	Cs	cesium	376	0.8	302	944	1.873	238
56	Ba	barium	503	0.9	1000.	2170.	3.62	206
57	La	lanthanum	538	1.1	1193	3737	6.15	194
Elements 58-71 have been omitted.								
72	Hf	hafnium	659	1.3	2506	4876	13.3	164
73	Ta	tantalum	728	1.5	3290.	5731	16.4	158
74	W	tungsten	759	1.7	3695	5828	19.3	150.
75	Re	rhenium	756	1.9	3458	5869	20.8	141
76	Os	osmium	814	2.2	3306	5285	22.587	136
77	Ir	iridium	865	2.2	2719	4701	22.562	132
78	Pt	platinum	864	2.2	2041	4098	21.5	130.
79	Au	gold	890.	2.4	1337	3129	19.3	130.
80	Hg	mercury	1007	1.9	234	630.	13.5336	132
81	Tl	thallium	589	1.8	577	1746	11.8	144
82	Pb	lead	716	1.8	600.	2022	11.3	145
83	Bi	bismuth	703	1.9	544	1837	9.79	150.
84	Po	polonium	812	2.0	527	1235	9.20	142
85	At	astatine	—	2.2	575	—	—	148
86	Rn	radon	1037	—	202	211	0.009074	146
87	Fr	francium	393	0.7	300.	—	—	242
88	Ra	radium	509	0.9	969	—	5	211
89	Ac	actinium	499	1.1	1323	3471	10.	201
Elements 90 and above have been omitted.								

*boiling point at standard pressure

** density of solids and liquids at room temperature and density of gases at 298 K and 101.3 kPa

Table T
Important Formulas and Equations

Density	$d = \frac{m}{V}$	d = density m = mass V = volume
Mole Calculations	number of moles = $\frac{\text{given mass}}{\text{gram-formula mass}}$	
Percent Error	$\% \text{ error} = \frac{\text{measured value} - \text{accepted value}}{\text{accepted value}} \times 100$	
Percent Composition	$\% \text{ composition by mass} = \frac{\text{mass of part}}{\text{mass of whole}} \times 100$	
Concentration	parts per million = $\frac{\text{mass of solute}}{\text{mass of solution}} \times 1\,000\,000$	
	molarity = $\frac{\text{moles of solute}}{\text{liter of solution}}$	
Combined Gas Law	$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$	P = pressure V = volume T = temperature
Titration	$M_A V_A = M_B V_B$	M_A = molarity of H^+ M_B = molarity of OH^- V_A = volume of acid V_B = volume of base
Heat	$q = mC\Delta T$ $q = mH_f$ $q = mH_v$	q = heat m = mass C = specific heat capacity ΔT = change in temperature H_f = heat of fusion H_v = heat of vaporization
Temperature	$K = ^\circ\text{C} + 273$	K = kelvin $^\circ\text{C}$ = degree Celsius

Prefix	Meaning	Meaning
Kilo	1000/1	1000m in 1km
L, m, or g (base unit)	1/1	
Centi	1/100	1m in 100cm
Milli	1/1000	1m in 1000mm

Non - SI unit conversions

Unit	Unit amount	SI equivalent
Pounds	2.2	1Kg
Cup	4.2	1 L
Oz.	.035	1 g
Feet	3.28	1 m
Oz.	33.8	1 L
mile	.00062137	1 m

TABLE V

TABLE A-22 STANDARD ELECTRODE POTENTIALS

Electrode reaction	E° (V)
$Li^+(aq) + e^- \rightleftharpoons Li(s)$	-3.0401
$K^+(aq) + e^- \rightleftharpoons K(s)$	-2.931
$Ca^{2+}(aq) + 2e^- \rightleftharpoons Ca(s)$	-2.868
$Na^+(aq) + e^- \rightleftharpoons Na(s)$	-2.71
$Mg^{2+}(aq) + 2e^- \rightleftharpoons Mg(s)$	-2.372
$Al^{3+}(aq) + 3e^- \rightleftharpoons Al(s)$	-1.662
$Zn(OH)_2(s) + 2e^- \rightleftharpoons Zn(s) + 2OH^-(aq)$	-1.249
$2H_2O(l) + 2e^- \rightleftharpoons H_2(g) + 2OH^-(aq)$	-0.828
$Zn^{2+}(aq) + 2e^- \rightleftharpoons Zn(s)$	-0.7618
$Fe^{2+}(aq) + 2e^- \rightleftharpoons Fe(s)$	-0.447
$PbSO_4(s) + H_3O^+(aq) + 2e^- \rightleftharpoons Pb(s) + HSO_4^-(aq) + H_2O(l)$	-0.42
$Cd^{2+}(aq) + 2e^- \rightleftharpoons Cd(s)$	-0.4030
$Pb^{2+}(aq) + 2e^- \rightleftharpoons Pb(s)$	-0.1262
$Fe^{3+}(aq) + 3e^- \rightleftharpoons Fe(s)$	-0.037
$2H_3O^+(aq) + 2e^- \rightleftharpoons H_2(g) + 2H_2O(l)$	0.000
$AgCl(s) + e^- \rightleftharpoons Ag(s) + Cl^-(aq)$	+0.222
$Cu^{2+}(aq) + 2e^- \rightleftharpoons Cu(s)$	+0.3419
$O_2(g) + 2H_2O(l) + 4e^- \rightleftharpoons 4OH^-(aq)$	+0.401
$I_2(s) + 2e^- \rightleftharpoons 2I^-(aq)$	+0.5355
$Fe^{3+}(aq) + e^- \rightleftharpoons Fe^{2+}(aq)$	+0.771
$Hg_2^{2+}(aq) + 2e^- \rightleftharpoons 2Hg(l)$	+0.7973
$Ag^+(aq) + e^- \rightleftharpoons Ag(s)$	+0.7996
$Br_2(l) + 2e^- \rightleftharpoons 2Br^-(aq)$	+1.066
$MnO_2(s) + 4H_3O^+(aq) + 2e^- \rightleftharpoons Mn^{2+}(aq) + 6H_2O(l)$	+1.224
$O_2(g) + 4H_3O^+(aq) + 4e^- \rightleftharpoons 6H_2O$	+1.229
$Cl_2(g) + 2e^- \rightleftharpoons 2Cl^-(aq)$	+1.358
$PbO_2(s) + 4H_3O^+(aq) + 2e^- \rightleftharpoons Pb^{2+}(aq) + 6H_2O(l)$	+1.455
$MnO_4^-(aq) + 8H_3O^+(aq) + 5e^- \rightleftharpoons Mn^{2+}(aq) + 12H_2O(l)$	+1.507
$PbO_2(s) + HSO_4^-(aq) + 3H_3O^+(aq) + 2e^- \rightleftharpoons PbSO_4(s) + 5H_2O(l)$	+1.691
$Ce^{4+}(aq) + e^- \rightleftharpoons Ce^{3+}(aq)$	+1.72
$Ag_2O_2(s) + 4H^+(aq) + e^- \rightleftharpoons 2Ag(s) + 2H_2O(l)$	+1.802
$F_2(g) + 2e^- \rightleftharpoons 2F^-(aq)$	+2.866

