

2. **Title = Skittles/isotope Lab** (you may write whatever you would like to about this lab in your journal. Remember that you can use your journal as a reference on your tests.)

Skittles Lab: Isotopes and Calculating Average Atomic Mass

Objective:

1. Show how Skittles can be an analogy for isotopes
2. Calculate the atomic mass for the element Sk (skittles)

Procedure:

1. Obtain a snack bag of skittles
2. Record the mass of each skittle as either 1.1, 1.0, or .9 grams (it would be easiest to make a data table of this information.)
3. Calculate what percent of each mass the skittles bag contains.
4. Multiply each percentage (in decimal form) by it's corresponding mass. (example = if 50% had a mass of .9g then multiply .5 by .9 and get an answer)
5. Calculate the average atomic mass of a skittle by adding the three answers in number 4.
6. Repeat this procedure for the class data.
7. Write \times bullets/sentences that discuss what was reinforced with this journal. ("beef up your journal")

Data: (individual)

skittle type (mass)	number of that type of skittle
.9	
1.0	
1.1	

Calculations = (individual)

1. % of skittles that are .9 = _____ contributing mass = _____

2. % of skittles that are 1.0 = _____ contributing mass = _____

3. % of skittles that are 1.1 = _____ contributing mass = _____

Average atomic mass of a skittle = _____

Data: (class)

skittle type (mass)	number of that type of skittle
.9	
1.0	
1.1	

Calculations = (individual)

4. % of skittles that are .9 = _____ contributing mass = _____

5. % of skittles that are 1.0 = _____ contributing mass = _____

6. % of skittles that are 1.1 = _____ contributing mass = _____

Average atomic mass of a skittle = _____